Chemistry (Math Notes)

Module #9

I. Key Terms/Concepts

Concentration:

Concentrated:

Diluted:

Strength:

Indicator:

Acid:

Base:

Amphiprotic:

II. Chemical Reactions

1. Formation:_____

2. Decomposition: _____

3. Combustion: (complete)_____

   Incomplete _____

4. Acid-Base _____

_____

 III. Molarity & Dilution Equation:

Concentration is amount / volume. Examples of concentration units are:
_____, _____, _____

In chemistry we often use _____ to measure concentration.
Molarity (M) is the number of moles / # liters of solution.

Chemists usually keep "stock" solutions, which are then diluted for use.

M_1 x V_1 = M_2 x V_2

M_1 is the molarity of the _____ _____.

V_1 is the molarity of the _____ _____.

M_2 is the molarity of the _____ _____that the chemist
wants.

V_2 is the volume of the _____ _____ that the chemist
wants.

(molarity) x (# liters) = #moles

IV. Titration

What is a titration?

It is _____. You always begin with a _____

_____ _____.

You take a known amount of acid & add a _____ to it slowly. An _____ will change color, indicating that you have added just enough base to eat up all the acid. This is the

_____, allowing you to determine the _____ of the acid. This technique also works in reverse, adding an acid slowly to a known amount of a base.