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2nd Period

Metal combustion Lab

RESEARCH QUESTION

What happens in the combustion of a metal?

INTRODUCTION

When heated, the metal will go through combustion. In the lab, we are going to carry out a combustion reaction using iron filling.

PROCEDURE

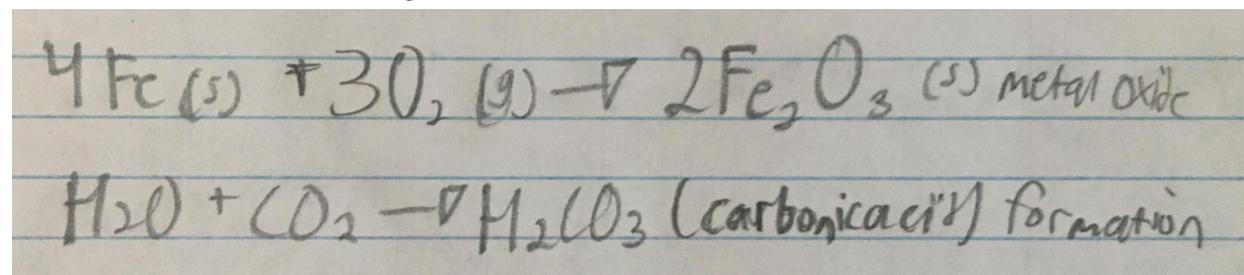
First we lit the burner. We then sprinkled the iron fillings into the flame. The iron fillings then turned into sparks and flew into the air.

RESULTS

When the sparks flew out and landed on the desk, we say that they had turned into something different. This new substance was rust. Rust is also known as ferric oxide. This material is darker than the fillings.

CONCLUSION

This Combustion of metal that happened was known as a chemical reaction. So we can conclude that the iron fillings turned into ferric oxide (rust). The fillings turned darker and more damaged compared to the iron fillings. In conclusion, the combustion of metals created a new substance, a metal oxide, being formed.



VOCABULARY:

Combustion reaction- A reaction in which a substance reacts with oxygen gas, releasing energy in the form of light and heat.

Metal-A solid material that is typically hard, shiny, and malleable, fusible, and ductile with good electrical and thermal conductivity.

Homonuclear diatomic-Consist of two atoms of the same element.