

Acid (A)	Base (B)
Lemons, vinegar, tomato juice	Bitter taste
Can burn; can be corrosive	Slippery to touch
Sour taste	Baking soda, soap, bleach

Fill in the blanks:

Atomic number	Salt	Ion	Atomic theory	Element
Chemistry	Atom	Chemical Change	Chemical formula	Chemical Symbol

- Based on the number of protons in the nucleus
- When scientists think atoms are like, based on repeated observations.
- A substance; only one kind of atom
- The smallest piece of an element that can be recognized as that element
- An atom that has gained or lost electrons
- Abbreviation for the name of an element
- The process in which different types of atoms join and form a new substance
- Chemical symbols and numbers abbreviating the name of the compound
- An ionic compound produced from a reaction between an acid and a base
- The study of matter-what it is made of, what its usual characteristics are, how it reacts to other matter.

True - False

- A neutral solution is neither acidic nor basic?
- In a stable atom, the outermost shell is completely closed or filled up?
- A reaction between an acid and a base can form salt?

Multiple Choice

4. The compound H_2CO_3 has _____ atom(s) of carbon in it.
- One
 - Two
 - Three
5. When positively and negatively charged ions attract each other they form a(n) _____?
- Ionic bond
 - Element
 - Covalent bond
6. A substance that changes color when exposed to acid or base solutions is called a(n) _____?
- Indicator
 - Salt
 - Molecule
7. The sum of the number of protons and neutrons results in an atom's approximate _____?
- Atomic number
 - Atomic mass
 - Electrons
8. You can determine the concentration of acids or bases in a solution by using _____?
- The periodic table of elements
 - A salt
 - The pH scale
9. _____ a scientist who developed a simplified model of an atom
- Robert Boyle
 - Dmitri Mendeleev
 - Niels Bohr
10. Dmitri Mendeleev developed the periodic table of elements, which _____ elements.
- Classifies
 - Weighs
 - Alphabetizes
11. A _____ reaction occurs when a compound is broken down into simpler compounds.
- Replacement
 - Synthesis
 - Decomposition