

Lab 5.2

Question – What behavior do ionic and covalent compounds exhibit when they are combined as a mixture?

Procedure – In book on page 196

Vocabulary -

Ionic Compound – A compound formed by the transfer of electrons between metals and non-metals.

Covalent Compound – A compound formed when two non-metals react with each other.

Polar Covalent Compound – A compound where two atoms share a pair of electrons unequally because of difference in their electro-negativities.

Non Polar (Purely) Covalent Compound – A compound where two atoms have the same electro-negativities,.

Mixture – A compound made up of two or more chemical components that are not chemically linked.

Immiscible – Liquids that wont mix to give a single phase.

Suspension – A heterogeneous mixture in which the solid particles do not dissolve, but get suspended throughout the bulk of the solvent, left floating around freely in the medium.

Data – Salt is ionic. It's chemical formula is NaCl, a metal and a non-metal chemically bonded. Water is a polar covalent, it carries a slight charge of attraction. Oil is purely covalent, it has no charges. There is no attractive force between the salt and the oil, so we see that it does not dissolve. The charges of the water and salt molecules interact so that they pull on each other. Thus we see salt dissolved in the water unlike the oil. Like compounds dissolves like.

Observation -

