

**Naming & Formula Writing for Type 1 Ionic Compounds**

What is a Type 1 Compound?

- These compounds are binary, in that they are made up of two types of elements.
- It is made up of a metal from columns 1, 2 or 13 and a nonmetal.
- These metals have only one charge or oxidation state.
- Group 1 metals have a +1 charge.
- Group 2 metals have a +2 charge.
- Group 13 metals have a +3 charge.

Steps for naming:

- Write the name of the metal.
- Write the root of the nonmetal and add the -ide suffix.

Examples of Naming:

- NaCl sodium chloride
- Al<sub>2</sub>S<sub>3</sub> aluminum sulfide

Correctly name the following compounds.

1. NaBr - *Sodium bromine*
2. Li<sub>2</sub>O - *lithium oxide*
3. NaCl - *Sodium chloride*
4. KI - *Potassium iodide*
5. CaS - *calcium sulfide*
6. MgO - *magnesium oxide*
7. CsF - *Cesium flouride*
8. AlCl<sub>3</sub> - *aluminum chloride*
9. MgI<sub>2</sub> - *magnesium iodide*
10. Rb<sub>2</sub>O - *rubidium oxide*
11. SrI<sub>2</sub> - *strontium iodide*
12. K<sub>2</sub>S - *potassium sulfide*