

## RESEARCH QUESTION

What is the difference between ionic and covalent compounds experimentally?

How can we break the bonds of compounds?

## INTRODUCTION

Ionic compounds don't have a definite shape and they have a high polarity. They also have high boiling points but stay solid at room temperature. Covalent compounds have a definite shape and a low polarity. They have a low boiling point and become gases at room temperature

## PROCEDURE

Pour approximately 150mL of H<sub>2</sub>O into each beaker. Then Put one rounded spatula of sugar into a beaker and mix thoroughly. After that put a different spatula with baking soda into the other beaker and mix well. Once you have finished mixing make sure your battery and leads are ready when its ready put one of the leads into your sugar beaker and observe. Then put the other lead from the same battery into your baking soda beaker and observe.

## OBSERVATIONS

Nothing happened when I put the battery lead into the sugar water. but in the other beaker the lead was fizzing and bubbling. And after a short while the tip of the lead turned to the color blue and even the water had a blue tint.

## CONCLUSION

Conclusion ionic compounds conduct electricity and covalents do not You can test using electrical currents. Baking soda is the ionic, sugar is the covalent. Ionics have metal in them, metals conduct electricity. The water overtime got a blueish tint because the lead was making blue bubbles.

Ionic: related to, composed of or using ions

Metal: a solid material that is typically hard, shiny or fusible with good electrical and thermal conductivity

Implication: the conclusion that can be drawn from something