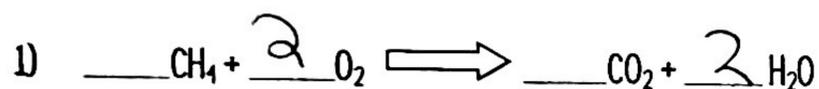


C H E M I S T R Y

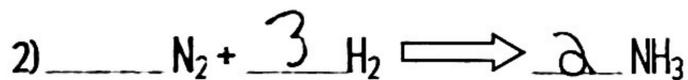
Balancing Equations: Counting Atoms

Guided Practice

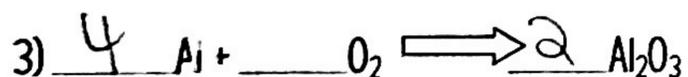
The Law of Conservation of Mass states that matter cannot be created nor destroyed, only rearranged. This means that all chemical equations must be balanced. To begin solving the equations on page, start first by counting the atoms on each side of the chemical equation. Then make the necessary changes to balance the equation.



ELEMENT	• REACTANTS	• PRODUCTS
CARBON	1	1
HYDROGEN	4	4
OXYGEN	4	4



ELEMENT	• REACTANTS	• PRODUCTS
NITROGEN	2	2
HYDROGEN	6	6



ELEMENT	• REACTANTS	• PRODUCTS
ALUMINUM	4	4
OXYGEN	6	6



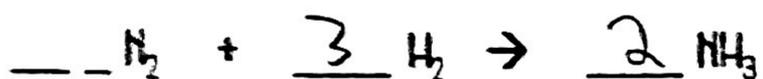
ELEMENT	• REACTANTS	• PRODUCTS
ARSENIC	1	1
SODIUM	1	3
OXYGEN	1	3
HYDROGEN	1	2

3. Reactants

N	2
H	3

Products

N	1
H	6



4. Reactants

K	1
Cl	2

Products

K	2
Cl	2



5. Reactants

C	1
F	4
Br	4

Products

C	1
F	4
Br	4



6. Reactants

Ga	1
F	3
Cs	3

Products

Ga	1
F	3
Cs	3

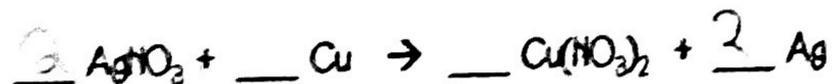


7. Reactants

Ag	2
N	2
O	6
Cu	1

Products

Ag	2
N	2
O	6
Cu	1



8. Reactants

Ba	
S	
Pt	
F	

Products

Ba	
S	
Pt	
F	

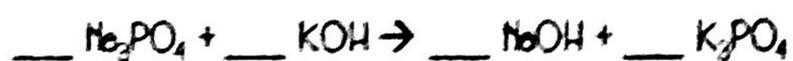


9. Reactants

Na	
P	
O	
K	
H	

Products

Na	
P	
O	
K	
H	



10. Reactants

Mg	
F	
Li	
C	
O	

Products

Mg	
F	
Li	
C	
O	

