

Moles Worksheet

- 1) Define "mole".
a standard scientific unit for measuring large quantities of very small entities. $6.02 \cdot 10^{23}$ of anything, usually atoms or molecules.
- 2) How many moles are present in 34 grams of $\text{Cu}(\text{OH})_2$?
0.35
- 3) How many moles are present in 2.45×10^{23} molecules of CH_4 ?
0.41
- 4) How many grams are there in 3.4×10^{24} molecules of NH_3 ?
96
- 5) How much does 4.2 moles of $\text{Ca}(\text{NO}_3)_2$ weigh?
689
- 6) What is the molar mass of MgO ?
40.3 / + grams / mole
- 7) How are the terms "molar mass" and "atomic mass" different from one another?
molar mass defines the mass of 1 mole of a compound and grams per mole and atomic mass defines one mole of an element.
- 8) Which is a better unit for expressing molar mass, "amu" or "grams/mole"?
grams/mole