Moles Worksheet

1) Define "mole".
6.02 x10^23 of anything but usually molecules or atoms

2) How many moles are present in 34 grams of $Cu(OH)_2$?
0.35 moles

3) How many moles are present in 2.45×10^{23} molecules of CH_4?
0.41 moles

4) How many grams are there in 3.4×10^{24} molecules of NH_3?
96 grams

5) How much does 4.2 moles of $Ca(NO_3)_2$ weigh?
689 grams

6) What is the molar mass of MgO?
40.3 grams/mole

7) How are the terms "molar mass" and "atomic mass" different from one another? Molar mass refers to the mass of one mole of a chemical compound, Atomic Mass refers to the mass of one mole of an element

8) Which is a better unit for expressing molar mass, "amu" or "grams/mole"?
Grams/mole