

Lecture Notes - Module 7

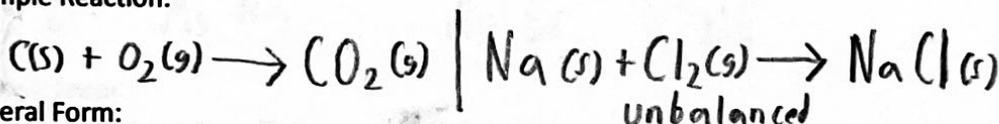
Youtube Video: Types of Chemical Reactions

Types of Chemical Equations:

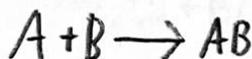
1. Synthesis
2. Decomposition
3. Combustion
4. Single replacement
5. Double replacement

A synthesis reaction is A compound made from simpler materials.

Example Reaction:



General Form:



A decomposition reaction is A compound broken down into simpler compounds, or all the way down to the elements that make it up.

Example Reaction

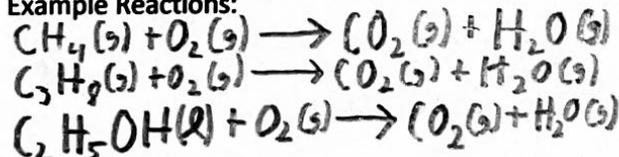


General Form

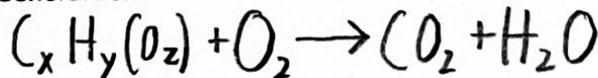


A combustion reaction is A compound containing carbon and hydrogen (and sometimes oxygen) combines with oxygen gas to produce carbon dioxide and water.

Example Reactions:

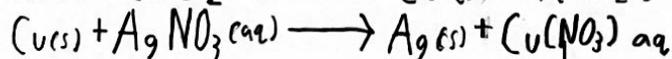
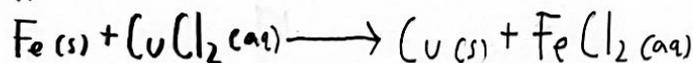
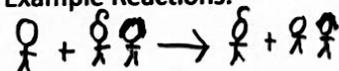


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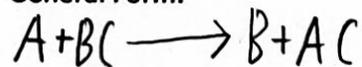


Single Replacement is One element that ~~starts out by itself~~ starts out by itself replaces another element in a compound, kicking it out.

Example Reactions:

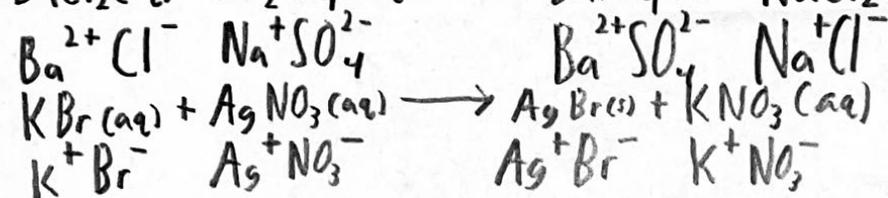
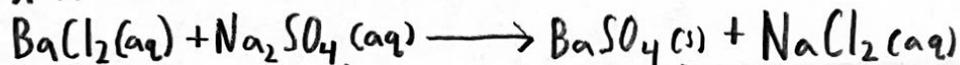
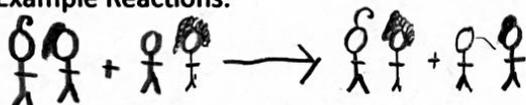


General Form:



Double Replacement is where the positive and negative ions in two compounds switch places.

Example Reactions:



General Form:

