

Lecture Notes - Module 7

Youtube Video: Types of Chemical Reactions

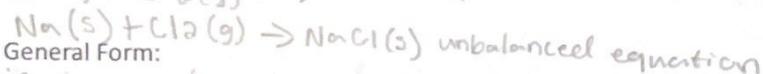
Types of Chemical Equations:

1. Synthesis
2. Decomposition
3. Combustion
4. Single Replacement
5. Double Replacement

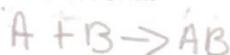
A synthesis reaction is making.

A compound is made from simpler materials

Example Reaction:



General Form:



A decomposition reaction is broken down.

A compound is broken down into simpler compounds, or all the way down to the elements that make it up.

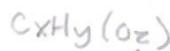


General Form



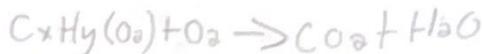
A combustion reaction is burning.

A compound containing carbon and hydrogen (sometimes oxygen) combines with oxygen gas to produce carbon dioxide and water.



This is the letter z, not a number 2. The number of oxygen also varies.

General Form:



Single Replacement is kicking out another one element.

one element that starts off by itself replaces another element in a compound, kicking it out

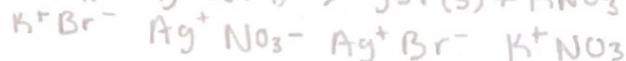
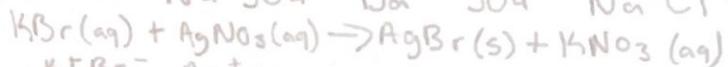


General Form:



Double Replacement is Positive / Negative ions. Nobody gets kicked out.
The positive and negative ions in two compounds switch places

Example Reactions:



General Form:

