

CHEM 150 - GENERAL CHEMISTRY II - LABORATORY REPORT

Experiment 10. Qualitative Analysis: Group I

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Section _____ Date Nov 29th, 2021 Instructor Dr. Grotak

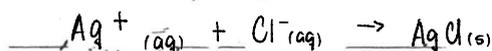
1. Unknown ~~number~~ Distilled water

2. Results of analysis of unknown:

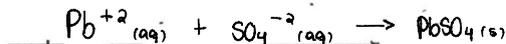
	Present	Absent
Silver ions	<u>X</u>	_____
Lead ions	_____	<u>X</u>
Mercurous ions	_____	<u>X</u>

3. Write net ionic equations for each of the following:

a. Reaction which caused silver ions to precipitate initially.



b. Confirmatory test for lead.



c. Reaction which caused the silver chloride precipitate to dissolve.



4. Suppose that a ~~Group I~~ ^{distilled water} unknown initially yielded a precipitate when HCl solution was added, but that the precipitate dissolved completely when hot water was added. In this case, indicate what ions were present in the unknown and which were absent.

If the unknown solution (distilled water) yielded a precipitate when HCl was added and the same precipitate was dissolved when hot water was added, this would mean that

Pb^{2+} ion was present as it forms a precipitate when reacting with HCl, however it dissolves when hot water is added.